

2.047 OH-, ozone-, and nitrate radicals-initiated aging of secondary organic aerosol formed from the ozonolysis of limonene.

Early Career Scientist

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Abstract:

Formation of secondary organic aerosol (SOA) from the reaction of VOCs with atmospheric oxidants (OH, O₃ and NO₃) and OH-initiated aging (OH reactions of secondary products) are included in the volatility-basis set model. However, night time aging by NO₃ and O₃ are not included. Significant nitrate aerosol formation was reported in recent field studies conducted in limonene abundant area, suggesting that the reactions of NO₃ with the secondary products formed from limonene oxidation as well as the direct reaction of NO₃ with limonene might contribute to nitrate aerosol formation. In a 6 m³ smog chamber we conducted light condition OH-initiated aging and dark condition O₃-, NO₃-initiated aging of limonene SOA to understand importance of daytime and night aging in SOA formation. Limonene SOA was produced from the ozonolysis of limonene. After SOA level became stable, OH-initiated aging was carried out by H₂O₂ photolysis and dark condition NO₃-, O₃- initiated aging were carried out by the injection of 1 ppm N₂ O₅ and O₃ respectively. Gas phase reactants and products were monitored by using FT-IR spectrometer and proton transfer reaction mass spectrometer. SOA formation was monitored by Aerodyne aerosol mass spectrometer (AMS) and SMPS. SOA particles formed before and during aging were collected on a Teflon filter and analysed by liquid chromatograph-mass spectrometer (LC-MS). AMS results showed that all three oxidants used in the experiments produced similar levels of SOA after one-hour aging, suggesting that NO₃- and O₃-initiated aging have comparable SOA formation potentials with OH-initiated aging. LC-MS results shows that ozonolysis of limonene produced monomeric and dimeric products in particle phase. Further, it shows that increasing of oligomeric products upon O₃ aging whereas several new monomeric and oligomeric compounds with higher mass range were produced upon NO₃ aging. This work was supported by JSPS KAKENHI Grant Numbers JP16H06305 and JP 17H01866.